



Thermodynamic properties and the behavior of H₂O in Fe- and Mg-cordierite

E. Dachs (1), C.A. Geiger (2)

1. Department of Material Science & Physics, University of Salzburg, Austria
(edgar.dachs@sbg.ac.at)
2. Institute for Geosciences, University of Kiel, Germany (chg@min.uni-kiel.de)

Cordierite, ideal formula $(\text{Mg,Fe})_2\text{Al}_4\text{Si}_5\text{O}_{18}\text{n}(\text{H}_2\text{O,CO}_2)$, is a key volatile-bearing phase in many metamorphic rocks. Thus, knowledge of its thermodynamic properties is necessary for a variety of investigations on the nature of fluids in Earth's crust. In order to enable more quantitative thermodynamic calculations, heat capacity measurements between 5 and 300 K were carried out on synthetic samples of anhydrous low Fe-cordierite (Fe-Cd), as well as on anhydrous and hydrous low Mg-cordierite (Mg-Cd), using heat-pulse calorimetry (HPC).

From our HPC results, the standard entropy values at 298.15 K for anhydrous Fe-Cd, and anhydrous and hydrous Mg-Cd were calculated as 460.5 ± 0.5 , 406.1 ± 0.4 and 450.9 ± 0.5 J/(mol·K), respectively. Heat capacity (C_p) polynomials for anhydrous and hydrous Fe-cordierite and anhydrous Mg-cordierite at $T > 270$ K are:

$$C_p^{\text{Fe-Cd}} = 911.1 (\pm 9.7) - 5829.2 (\pm 363) \cdot T^{-0.5} - 13.9424 (\pm 2.522) \cdot 10^6 \cdot T^{-2} + 1470.4 (\pm 454.84) \cdot 10^6 \cdot T^{-3},$$

$$C_p^{\text{hFe-Cd}} = 967.3 (\pm 9.7) - 6070.4 (\pm 363) \cdot T^{-0.5} - 13.9389 (\pm 2.522) \cdot 10^6 \cdot T^{-2} + 1470.4 (\pm 454.84) \cdot 10^6 \cdot T^{-3},$$

$$C_p^{\text{Mg-Cd}} = 882.0 (\pm 4.9) - 5155.8 (\pm 167) \cdot T^{-0.5} - 20.7584 (\pm 0.806) \cdot 10^6 \cdot T^{-2} + 2736.0 (\pm 112.73) \cdot 10^6 \cdot T^{-3},$$

as derived from the HPC data and published DSC data. Based on the data of Paukov et al. (2007), a C_p -polynomial for H₂O in hydrous Mg-cordierite was derived as:

$$C_{p,H_2O}^{hMg-Cd} = 56.17 (\pm 0.44) - 241.2 (\pm 4.5) \cdot T^{-0.5} + 3473.5 (\pm 272.1) \cdot T^{-2} - 11658.7 (\pm 1420.3) \cdot T^{-3}.$$

Using phase equilibrium data for the reaction $3 \text{ Fe-Cd-H}_2\text{O} = 2 \text{ almandine} + 4 \text{ silimanite} + 5 \text{ quartz} + 3 \text{ H}_2\text{O}$ (Mukhopadhyay & Holdaway 1994), $\Delta_f H^\circ = -8448.26$ kJ/mol was obtained for anhydrous Fe-Cd and the values $\Delta_f H^\circ = -8750.23$ kJ/mol and $S^\circ = 520.6$ J/(mol·K) were derived for hydrous Fe-Cd. Phase relations in the FeO-Al₂O₃-SiO₂-H₂O system and isohydrans for H₂O in Fe-Cd were calculated and the phase diagrams constructed.

References

- Mukhopadhyay B., Holdaway M.J. (1994) *Contrib. Mineral. Petrol.*, **116**, 462-472.
- Paukov I.E., Kovalevskaya Yu.A., Rahmoun N.S., Geiger C.A. (2007): *Am. Mineral.*, **92**, 388-396.